Periodic Classification Of Elements Guess Question X Chemistry

Q1 In which form matter is present around us?

Q2 Define element.

Q3 At present, how many elements are known to us?

Q4 The earliest attempt in classifying elements resulted in the formation of two groups of elements. What are they?

Q5 Who made the first attempt of classifying elements?

Q6 On what basis Döbereiner classified elements?

Q7 Döbereiner classified elements into how many groups?

Q8 What name was given to Döbereiner groups?

Q9 What is the total number of elements in Döbereiner groups?

Q10 How did John Newlands classify elements?

Q11 Name the first element of Newland's octaves.

Q12 Name the last element of Newland's octaves.

Q13 What is your observation from Newland's octaves?

Q14 What is Newland's Law of octaves?

Q15 Which law shows that Beryllium and Magnesium resemble each other and How?

Q16 Newland's Law of octaves is applicable up to which element?

Q17 Newland's Law of octaves worked well with which type of elements?

Q18 Who deserves the main credit of classifying elements?

Q19 What is the fundamental property of elements?

Q20 Besides atomic masses, on what other basis were the elements arranged in the Mendleev's periodic table?

Q21 Which chemical property of an element was treated as one of the basic property for classifying elements and why?

Q22 What is Mendleev's periodic Law?

Q23 What name is given to vertical columns in Mendleev's periodic table?

Q24 What name is given to horizontal rows in Mendleev's periodic table?

Q25 While developing the Periodic table, at few places Mendleev inverted the sequence of some elements i.e. he placed an element with slightly greater atomic mass before the element of lower atomic mass. Why did he do so?

Q26 Though the atomic mass of cobalt (58.9) is greater than nickel (58.7) yet Co is placed before Ni in Mendleev's periodic table. Why?

Q27 Which elements did not exist at the time of Mendleev's periodic classification? What name was given to these elements?

Q28 In what way hydrogen resembles alkali metals?

Q29 In what way hydrogen resembles halogens?

Q30 Why hydrogen cannot be given a fixed position in periodic table?

Q31 What is the first limitation of Mendleev's periodic table?

Q32 What are isotopes?

Q33 How isotopes of all the elements posed a challenge to Mendleev's periodic table?

Q34 Who proposed that atomic number is the more fundamental property for classifying elements?

Q35 What is Modern periodic Law?

Q36 In Modern periodic table, How do elements belonging to the same group resemble each other? Write two points.

Q37 Different elements have same number of shells, in group or in period?

Q38 First period of the Modern periodic table contains only two elements. Justify.

Q39 How many elements are present in second group of the periodic table? Justify.

Q40 "The valence electrons determine the kind and number of bonds formed by an element". Justify.

Q41 An element belongs to the first group and third period of the periodic table. What conclusion can you draw from its position ?

Q42 What is atomic size?

Q43 What is valency?

Q44 A metal M forms an oxide having the formula M_2O_3. It belongs to the third period and thirteenth group of the Modern periodic table. Write the atomic number and valency of the element.

Q45 What were the two major shortcomings of Mendleev's periodic table? How have these been removed in the modern periodic table?

Q46 Two elements X and Y have atomic numbers 12 and 16 respectively. Write the electronic configuration for these elements. To which period of the modern periodic table do these two elements belong? What type of bond will be formed between them and Why?

Q47 What were the two achievements of Mendleev's periodic table? What was the basis of classification of elements in it?

Q48 Atomic radius decreases in moving from left to right in a period. Why?

Q49 Atomic radius increases down the group. Why?

Q50 In the modern periodic table a zig-zag line separates metals from non-metals. What are these elements called and why?